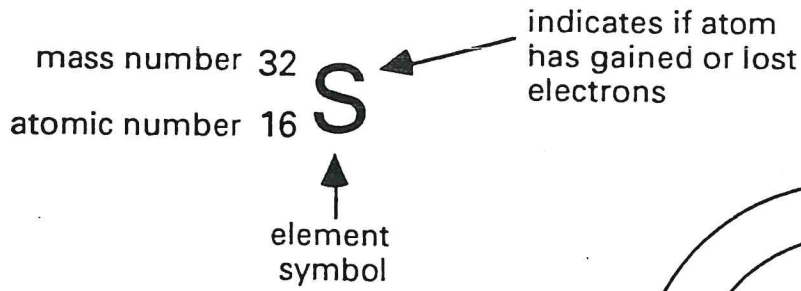
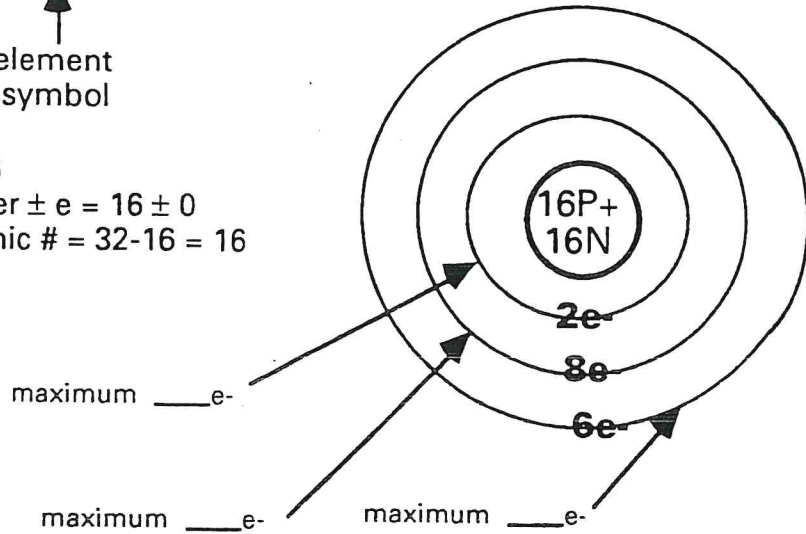


BOHR-ORBITAL DIAGRAMS



$\#P^+ = \text{atomic \#} = 16$
 $\#e^- = \text{atomic number} \pm e = 16 \pm 0$
 $\#N = \text{mass \#} - \text{atomic \#} = 32 - 16 = 16$



atomic number : _____

mass number: _____

ion: _____

cation: _____

anion: _____

how does an atom gain a positive charge? _____

how does an atom gain a negative charge? _____

Draw the BOHR ORBITAL DIAGRAM for the following:



Element Name	#P, #e, #N	Bohr Diagram	Lewis Structure
Hydrogen Atomic#: Mass:	#P: #e: #N:		
Helium Atomic#: Mass:	#P: #e: #N:		
Lithium Atomic#: Mass:	#P: #e: #N:		
Beryllium Atomic#: Mass:	#P: #e: #N:		
Boron Atomic#: Mass:	#P: #e: #N:		
Carbon Atomic#: Mass:	#P: #e: #N:		
Nitrogen Atomic#: Mass:	#P: #e: #N:		